

Jadavpur University
Department of Chemistry
1st year 2nd semester (CBCS system)
INORGANIC CHEMISTRY PRACTICAL

Paper: Chem/PR/03

Unit: 203L-I: Inorganic Chemistry-I

1. Estimate the amount of Fe(II) in the supplied sample using supplied standard $K_2Cr_2O_7$ solution. 40
2. Lab note book 5
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Procedure

1. Transfer the supplied sample **marked "I"** into a 250 ml volumetric flask quantitatively and make up the volume with distilled water. Consider this as the 'stock solution'.
2. Take an aliquot of 25 ml of the stock solution in a 250 ml conical flask, dilute with an equal volume of 4(N) sulphuric acid and add 5 ml syrupy phosphoric acid and 2-3 drops of BDS indicator solution. Titrate with the supplied standard $K_2Cr_2O_7$ solution from a burette, while swirling the flask gently until the solution changes from bright green to intense red-violet. Record the titre value of the dichromate solution from the burette. Repeat the titration thrice and calculate the strength of the supplied Fe(II) solution.
3. Report amount iron present in the supplied solution in the unit of g/L.